4. lodide ions can be oxidised using acidified potassium permanganate solution.

The equations are

$$2I^{-}(aq) \rightarrow I_{2}(aq) + 2e^{-}$$
  
 $MnO_{4}^{-}(aq) + 8H^{+}(aq) + 5e^{-} \rightarrow Mn^{2+}(aq) + 4H_{2}O(\ell)$ 

How many moles of iodide ions are oxidised by one mole of permanganate ions?

- A 1
- 3 2
- C 5
- D 10

